

Quantum Mechanics

- Predicts available energy states agreeing with Bohr.
- Don't have definite electron position, only a probability function.
- Each orbital can have 0 angular momentum!
- Each electron state labeled by 4 numbers:

n = principal quantum number (1, 2, 3, ...)

ℓ = angular momentum (0, 1, 2, ... $n-1$)

m_ℓ = component of ℓ ($-\ell < m_\ell < \ell$)

m_s = spin ($-1/2$, $+1/2$)

Coming Soon!