Arrhenius Equation

T dependence of a rate constant, k

 $\mathbf{k} = z p e^{-\mathbf{E}a/\mathbf{RT}}$

- $E_a = activation energy (kJ/mol)$
- $R = gas constant (8.314 x 10^{-3} kJ/K mol)$
- T = temperature (K)
- z = collision frequency
- p = steric factor (<1)