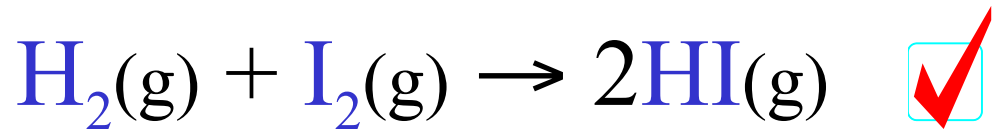
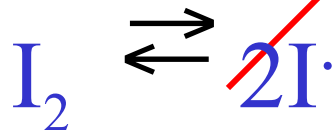


Reaction Mechanism

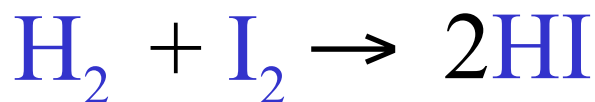


$$\text{rate} = k [\text{H}_2] [\text{I}_2]$$

step 1



step 2



need H_2 in the rate determining step

from *step 2* $\text{rate} = k [\text{H}_2] [\text{I}\cdot]^2$

$\text{I}\cdot$ = intermediate