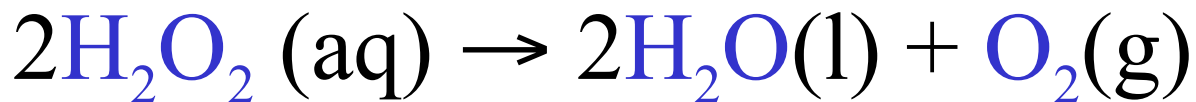


Reaction mechanism



$$\Delta G^\circ_{\text{rxn}} = [2(-237.9)] - [2(-131.67)] = -212.46 \text{ kJ}$$

spontaneous reaction

experimental rate law: $\text{rate} = k[\text{H}_2\text{O}_2][\text{I}^-]$

I^- = catalyst

increase rate of reaction

not consumed in the overall reaction

reactant in early elementary step

product in later elementary step