

1st order reactions

$$\ln \left(\frac{[A]_t}{[A]_0} \right) = -kt$$

$$\ln[A]_t = -kt + \ln[A]_0$$



$$\text{rate (M s}^{-1}\text{)} = k [\text{N}_2\text{O}_5(\text{M})] \quad k = 5.1 \times 10^{-4} \text{ s}^{-1}$$