

Pseudo 1st order reactions

high order reactions difficult to analyze

put in large excess of all but one reagent

rate = k [A] ^a [B] ^b	[A] ₀	[B] ₀
[B] ≈ constant	1.0 x 10 ⁻³ M	1.0 M
	-0.5 x 10 ⁻³ mol	-0.5 x 10 ⁻³ mol
rate = k' [A] ^a	0.5 x 10 ⁻³ M	0.999 M

$$k[A]^a[B]^b = k'[A] \quad k = \frac{k'}{[B]^b}$$