Integrated rate laws

second order reactions   rate = k $[A]^2$

$$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$$

many second order reactions rate = k$[A][B]$

A + B → C   A and B consumed stoichiometrically

$[A]_0 = [B]_0$

if not, no analytical solution