

Electrolysis of water



current
and time

charge

mol e⁻

mol
product

gram
product

$$(C/s) \times s \times \frac{\text{mol e}^-}{\text{C}} \times \frac{\text{mol O}_2}{\text{mol e}^-} \times \frac{\text{g O}_2}{\text{mol O}_2} = \text{g O}_2$$

↓ ↓ ↓ ↓

2.5 A $\frac{1 \text{ mol e}^-}{96500 \text{ C}}$ $\frac{1}{4}$ 32.0 g/mol 3.2