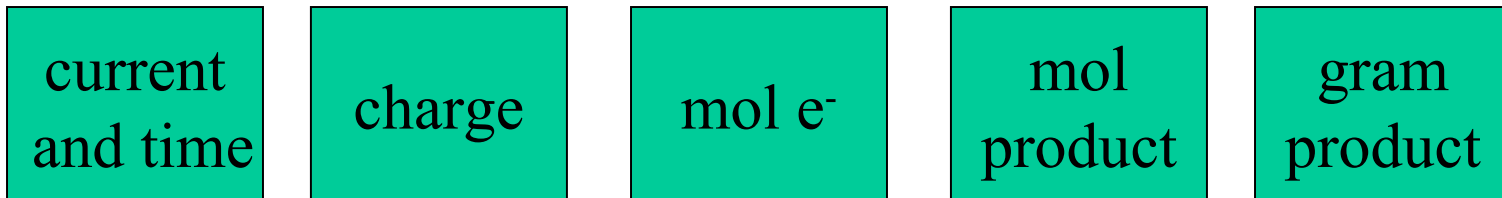
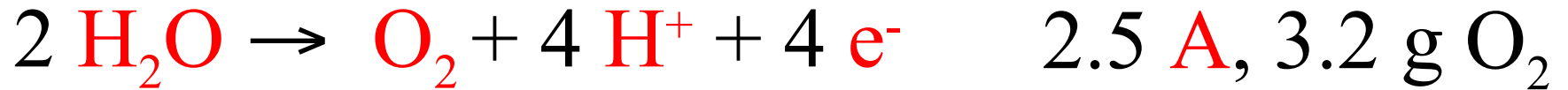


Electrolysis of water



$$\begin{array}{cccccc}
 (\text{C/s}) \times \text{s} \times \frac{\text{mol e}^-}{\text{C}} \times \frac{\text{mol O}_2}{\text{mol e}^-} \times \frac{\text{g O}_2}{\text{mol O}_2} = \text{g O}_2 \\
 \downarrow \qquad \qquad \downarrow \qquad \qquad \downarrow \qquad \qquad \downarrow \qquad \qquad \downarrow \\
 2.5 \text{ A} \qquad \frac{1 \text{ mol e}^-}{96500 \text{ C}} \qquad \frac{1}{4} \qquad 32.0 \text{ g/mol} \qquad 3.2
 \end{array}$$