Electrolysis of water

2.5 amp Power source → 3.2 g O_2

current = amperes (A) = coulombs/sec (C/s)

| current and time | charge | mol e$^-$ | mol product | gram product |

$$A(C/s) \times s \times \frac{1 \text{mol e}^-}{96,500 \text{ C}} \times \frac{\text{mol product}}{\text{mol e-}} \times \frac{\text{g product}}{\text{mol product}}$$