Electroplating

$0.75(\underline{C}) \text{ x } 25 \text{ min x } \underline{60}\text{s} = 1.125 \text{ x } 10^3 \text{ C}$

      s   min

$1.125 \text{ x } 10^3 \text{ C x } \dfrac{1 \text{ mol e}^-}{96500 \text{ C}} = 1.17 \text{ x } 10^{-2} \text{ mol e}^-$

$1.17 \text{ x } 10^{-2} \text{ mol e}^- \text{ x } \dfrac{1\underline{\text{mol Cu}}}{2\text{mol e}^-} = 5.83 \text{ x } 10^{-3}\text{mol Cu}$

$$\dfrac{0.37 \text{ g Cu}}{5.8\text{x}10^{-3} \text{ mol Cu}} = \dfrac{63.5 \text{ g}}{\text{mol}} \text{ atomic mass Cu}$$