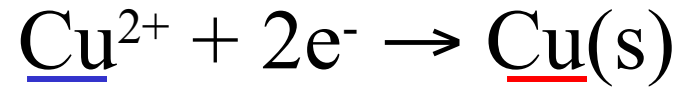


Concentration Cells



oxidation



reduction

$$Q = \frac{[\text{products}]_{\text{initial}}^m}{[\text{reactants}]_{\text{initial}}^n} = \frac{[\text{Cu}^{2+}]}{[\text{Cu}^{2+}]} = \frac{0.5}{2} = 0.25$$

$$\xi_{\text{cell}} = \xi_{\text{cell}}^{\circ} - \frac{RT}{nF} \ln Q = 0 - \frac{RT}{2F} \ln .25 = .02 \text{ V}$$