

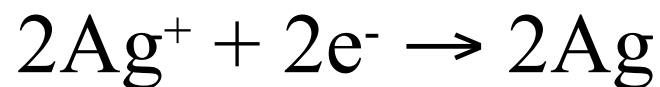
Nernst Equation

$$\xi_{\text{cell}} = \xi_{\text{cell}}^{\circ} - \frac{RT}{nF} \ln Q$$

anode

nF

cathode



$$\xi_{\text{cell}} = \underline{0.46\text{V}} - \frac{(8.314)(298) \ln 4}{2(96,500)} = \underline{0.44\text{V}}$$