

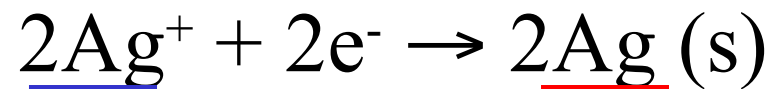
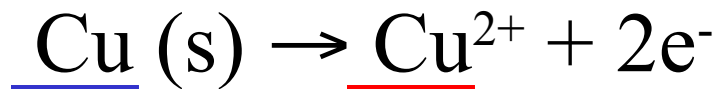
# Nernst Equation

$$\xi_{\text{cell}} = \xi_{\text{cell}}^{\circ} - \frac{RT}{nF} \ln Q$$

anode

nF

cathode



$$Q = \frac{[\text{products}]^m_{\text{initial}}}{[\text{reactants}]^n_{\text{initial}}} = \frac{[\text{Cu}^{2+}]}{[\text{Ag}^{+}]^2} = \frac{4}{1^2} = 4$$