

Nernst Equation

$$\xi_{\text{cell}} = \xi_{\text{cell}}^{\circ} - \frac{RT}{nF} \ln Q$$

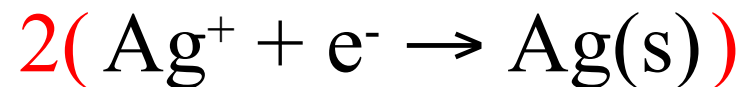
anode

nF

cathode



$$\xi^{\circ} = 0.34 \text{ V}$$



$$\xi^{\circ} = 0.80 \text{ V}$$

$$\xi_{\text{cell}}^{\circ} = \xi_{\text{red}}^{\circ} - \xi_{\text{ox}}^{\circ} = 0.80 - 0.34 = 0.46 \text{ V}$$