ξ^o_{cell}

$$\square^o_{cell} = \square^o_{red} - \square^o_{ox}$$

$Ag^+ (aq) + e^- \square Ag(s)$ $\square^o = 0.80$ V

$Al^{3+} (aq) + 3e^- \square Al(s)$ $\square^o = -1.66$ V

$$\square^o_{cell} = 0.80 - (-1.66) = 2.46 \text{ V}$$

$\square^o_{cell} > 0$ voltaic or galvanic cell spontaneous

$\square^o_{cell} < 0$ electrolytic cell non-spontaneous