Reduction Potential Al(s) $\Box Al^{3+}(aq) (1M) \Box \Box Ag^{+}(aq) (1M) \Box Ag(s)$ Al(s) \Box Al³⁺ + 3e⁻ $Ag^+ + e^- \Box Ag$ $3Ag^+ + Al \square 3Ag + Al^{3+}$ $\Box G^{o} = -712 \text{ kJ/mol}$ $\mathbf{W}_{\max} = -n F \square^{o}$ $\Box G^{o} = W_{max}$ \Box° = standard reduction potential standard = 1M, 1 atm reduction potential = tendency to gain e⁻