$$pH = pK_a + \log \frac{[A^-]}{[HA]}$$

$K_b = 1.8 \times 10^{-3}$

What is the pH of 100 mL of a 0.10 M solution of NH_3

a) 1.37   b) 1.87   c) 10.45   d) 12.13   e) 12.63

What volume of 1.00 M HCl is needed to reach the equivalence point?

a) 1000 mL   b) 100 mL   c) 10 mL   d) 1 mL   e) 0.1 mL

What is the pH at the equivalence point?

a) 2.35   b) 2.87   c) 3.25   d) 6.15   e) 11.65

What is the pH at the half-way point?

a) 2.75   b) 6.15   c) 7.00   d) 10.35   e) 11.25