CH$_3$COOH   NaOH

1.00 M   0.375 mol H$^+$ = 0.375 mol OH$^-$
                                              1.00 M

500 mL   add 375 mL

0.500 mol - 0.375 mol = .125 mol CH$_3$COOH = 0.143 M
                                      $\overline{\hspace{4cm}}$
                                            0.875 L

$\dfrac{0.375\ mol}{0.875\ L}$ = 0.429 M CH$_3$COO$^-$

Henderson-Hasselbalch Equation

$$pH = pK_a + \log \frac{[A^-]}{[HA]}$$

$$pH = 4.74 + \log \frac{.429}{.143} = 5.22$$