

Salts

product of neutralization reaction

strong **base**

weak **acid**

1.00 M **NaOH**

150 mL 0.500 M **HF**

$$\text{mol OH}^- = \text{mol H}^+$$

$$\frac{x \text{ L}}{\text{L}} \left| \frac{1.00 \text{ mol}}{\text{L}} = 0.075 \text{ mol} = \frac{0.15 \text{ L}}{\text{L}} \right| \frac{0.500 \text{ mol}}{\text{L}}$$
$$x = 0.075 \text{ L}$$



$$K_a = 1.5 \times 10^{-11}$$

NaF (s)