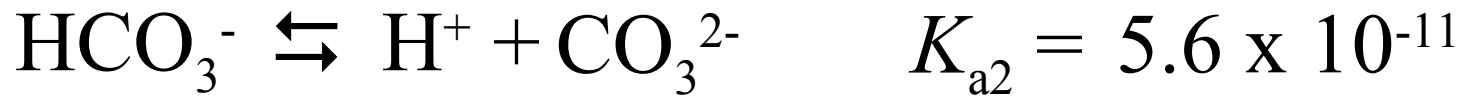


Find the pH of a **0.0037** M solution of  $\text{H}_2\text{CO}_3$



$$4.3 \times 10^{-7} = \frac{x^2}{3.7 \times 10^{-3}} \quad x = \frac{4.0 \times 10^{-5}}{3.7 \times 10^{-3}} \times 100 = 1.10 \%$$

$$\text{pH} = 4.40$$



<b>I</b>	$3.7 \times 10^{-3}$	0.00	0.00	
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$$[\text{CO}_3^{2-}] =$$



	$3.7 \times 10^{-3}$	$4.0 \times 10^{-5}$	$4.0 \times 10^{-5}$	
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