pH calculations

strong bases   complete dissociation

What is the pH of a 0.011 M solution of $Ca(OH)_2$?

$$Ca(OH)_2 + H_2O \rightarrow Ca^{2+} (aq) + 2(OH^-) (aq)$$
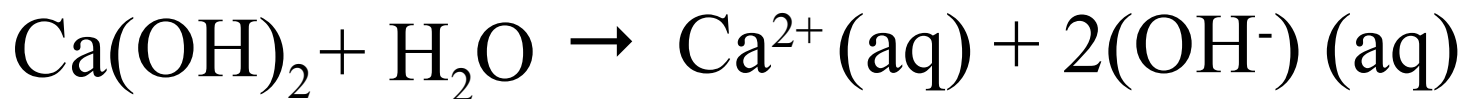
$[OH^-] = 2 [Ca^{2+}] = 0.022$ M $pOH = - \log [OH^-]$

$pH + pOH = 14.00$ $pOH = 1.67$

$pH = 12.33$

$$Na_2O + H_2O \rightarrow 2 Na^+ (aq) + 2 OH^- (aq)$$