$$\overset{+3}{C_2O_4{}^{2-}} \rightarrow \overset{+4}{2\,CO_2} + 2e^-$$

This is a balanced half-reaction

electrons are a product

$C_2O_4{}^{2-}$ is losing electrons

$C_2O_4{}^{2-}$ is being oxidized

Now combine the two half-reactions