

pH



$pH = -\log [H^+]$



low pH -1 15 high pH

acidic

basic

neutral $[H^+] = [OH^-] = 1.0 \times 10^{-7}$ pH = 7

$pH + pOH = 14$

$pOH = -\log [OH^-]$

pH of milk = 6.4 $[H^+] = 3.981071706 \times 10^{-7}$

$10^{-6.4} \quad 4 \times 10^{-7}$



1 sig. fig.