$$2H_2O + Al(s) + MnO_4 \rightarrow Al(OH)_4 + MnO_2$$

 H_2O What is being oxidized?

What is oxidation?

LEO

How many e⁻?

What is being reduced?

GER

What is the reducing agent?

A1

 MnO_{4}

 $Al(OH)_{4}$

 MnO_2