Clicker Question

Given the following data

$$2NH_3(g) \rightarrow N_2(g) + 3H_2(g) \qquad \Delta H = 92kJ$$
$$2H_2(g) + O_2(g) \rightarrow 2H_2O(g) \qquad \Delta H = -484kJ$$

Determine ΔH for

$$2N_2(g) + 6H_2O(g) \rightarrow 3O_2(g) + 4NH_3(g)$$
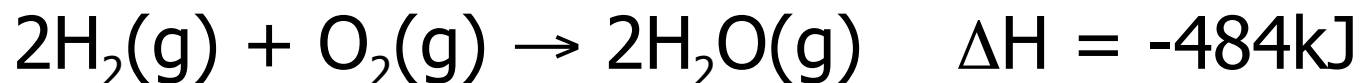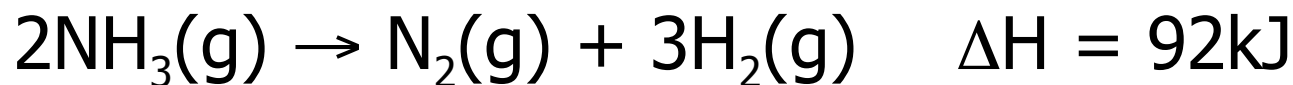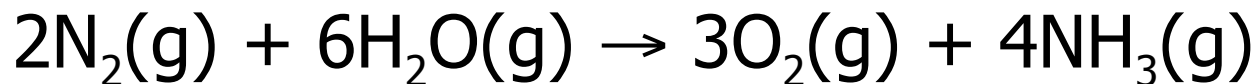
a) -1268 kJ
b)   1268 kJ
c) 1636 kJ
d) -1636 kJ
e)   I did not/cannot get any of these answers.