## Standard Enthalpy of Formation: $\Delta H_f^{\circ}$

Change in enthalpy that accompanies the formation of 1 mole of a compound from its elements with all substances in their standard states.

## Example:

$$H_2(g) + (1/2)O_2(g) \rightarrow H_2O(g) \quad \Delta H_f^{\circ} = -242kJ$$